

Chapter 17: Chemical Equilibrium

Worksheet 17.1: Equilibrium Constants

1. For the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$, $K_c = 5.0$ at 300 K.

- Calculate K_c for the reaction $\text{NO}_2(g) \rightleftharpoons \frac{1}{2}\text{N}_2\text{O}_4(g)$ at 300 K.
- Calculate K_c for the reaction $\text{N}_2\text{O}_4(g) \rightleftharpoons 2\text{NO}_2(g)$ at 300 K.
- Calculate K_c for the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2(g) + \text{O}_2(g)$ at 300 K.
- Calculate K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}_2(g)$ at 300 K.
- Calculate K_c for the reaction $\text{N}_2(g) + 2\text{O}_2(g) \rightleftharpoons 2\text{NO}_2(g)$ at 300 K.
- Calculate K_c for the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2(g) + 2\text{O}_2(g)$ at 300 K.
- Calculate K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ at 300 K.
- Calculate K_c for the reaction $2\text{NO}(g) \rightleftharpoons \text{N}_2(g) + \text{O}_2(g)$ at 300 K.
- Calculate K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ at 300 K.
- Calculate K_c for the reaction $2\text{NO}(g) \rightleftharpoons \text{N}_2(g) + \text{O}_2(g)$ at 300 K.

Answers

- 0.20
- 0.050
- 0.050
- 0.050
- 0.050
- 0.050
- 0.050
- 0.050
- 0.050
- 0.050

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