

## Stoichiometry Review

1. The combustion of propane occurs via the reaction:



How many grams of oxygen are required to burn completely 10.0 g of propane?

2. Small quantities of chlorine can be prepared in the laboratory by the reaction:



What volume of chlorine at STP can be prepared from 100.0 g of manganese dioxide?

3. The most common ore of arsenic is mispickel,  $\text{FeSAs}$ . Upon heating this ore, free arsenic is obtained. How many grams of  $\text{FeSAs}$  are required to produce 10.0 g of arsenic according to the following reaction?



4. Glucose is used as an energy source by the human body. The overall reaction in the body is:



Calculate the volume of oxygen at STP required to convert 28.0 g of glucose to carbon dioxide and water.

5. Potassium nitrate is widely used as a fertilizer because it provides two essential elements, potassium and nitrogen. It is made by mixing potassium chloride and nitric acid in the presence of oxygen according to the equation



How many kilograms of potassium nitrate will be produced from 50.0 kg of potassium chloride and 50.0 kg of nitric acid with excess oxygen?

6. Bromine can be prepared by adding chlorine to an aqueous solution of sodium bromide. How many grams of bromine are formed if 25.0 g of chlorine and 25.0 g of sodium bromide are reacted according to the reaction below? Which reactant is in excess and how much remains after the reaction is complete?



7. How many millilitres of 2.00 M hydrochloric acid are required to react with 2.55 g of zinc according to the following reaction:



8. A 0.473 g sample of phosphorus is reacted with an excess of chlorine. Calculate the mass of phosphorus pentachloride produced.

9. How many grams of oxygen can be prepared by the decomposition of 25.0 g of mercury(II) oxide?

10. What volume of oxygen gas at STP can be prepared by the decomposition of 25.0 g of potassium chlorate?

11. How many grams of zinc are required for the replacement of 0.100 g of hydrogen from sulfuric acid?

12. What mass of magnesium chloride is required to react with 10.0 g of silver nitrate?

# Chapter 9 Review Stoichiometry Section 2 Work

**Lei Shi**

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red-to-white gradient semi-circle on the right side of the bar.

## Chapter 9 Review Stoichiometry Section 2 Work:

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**Handbook of Thin Film Materials: Deposition and processing of thin films** Hari Singh Nalwa, 2002 **Alkane C-H Activation by Single-Site Metal Catalysis** Pedro J. Pérez, 2012-09-15 Over the past decade much research effort has been devoted to the design and synthesis of new reagents and catalysts that can influence carbon hydrogen bond activation mainly because of the prospect that C H activation could enable the conversion of cheap and abundant alkanes into valuable functionalized organic compounds Alkane C H Activation by Single Site Metal Catalysis presents the current state of the art development in the catalytic systems for the catalytic transformations of alkanes under homogeneous conditions Chapter 1 offers a comprehensive summary of the main discoveries realized so far Chapter 2 reviews the so called electrophilic activation initiated by Shulp n in the late 60s and the base for the Catalytica system Chapter 3 examines the catalytic borylation of alkanes discovered by Hartwig whereas chapter 4 provides an updated vision of the alkane dehydrogenation reaction Chapter 5 covers the oxygenation of C H bonds a field of enormous interest with bioinorganic implications and finally chapter 6 presents the functionalization of alkane C H bonds by carbene or nitrene insertion The history of C H bond activation and the current research described in this book highlight the current research and present the reader with an outlook of this field which continues to be explored by an increasingly visionary and enthusiastic group of organic organometallic biological and physical chemists *Math Review Toolkit for General Chemistry* Gary Long, 2006-05-30

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