

## Chapter 9 – Stoichiometry – Chapter Review

Last page worked out for you.

Match each description on the right to the correct term on the left.

- |          |  |                      |
|----------|--|----------------------|
| <u>B</u> | 1. The substance that determines the amount of product that can be formed in a reaction. | A. Percent yield     |
| <u>E</u> | 2. The amount of product that forms when a reaction is carried out in the laboratory.    | B. Limiting reagent  |
| <u>D</u> | 3. The calculation of quantities in chemical equations.                                  | C. Theoretical yield |
| <u>A</u> | 4. The ratio of the actual yield to the theoretical yield expressed as a percent.        | D. Stoichiometry     |
| <u>F</u> | 5. The substance that is present in enough quantity to react with a limiting reagent.    | E. Actual yield      |
| <u>C</u> | 6. The maximum amount of products that could be formed from given amounts of reactants.  | F. Excess reagent    |

Multiple Choice - Choose the best answer and write its letter in the blank.

- D 7. How many moles of  $\text{H}_2\text{O}$  would be required to produce 2.5 moles of  $\text{O}_2$ ?
- $$2 \text{H}_2\text{O} \rightarrow 2 \text{H}_2 + \text{O}_2$$
- A. 2.0 mol                      C. 4.0 mol  
B. 2.5 mol                      D. 5.0 mol
- D 8. If 3.00 mol of  $\text{CaCO}_3$  undergo decomposition to form  $\text{CaO}$  and  $\text{CO}_2$ , how many grams of  $\text{CO}_2$  are produced?
- $$\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$$
- A. 3.00 g                      C. 88.0 g  
B. 44.0 g                      D. 132 g
- B 9. How many moles of  $\text{H}_2\text{O}$  are produced when 240 grams of  $\text{CuO}$  react?
- $$\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$$
- A. 1.0 mol                      C. 18 mol  
B. 3.0 mol                      D. 54 mol

# Chapter 9 Stoichiometry Multiple Choice Answers

**Neil D. Jespersen, Pamela Kerrigan**



## Chapter 9 Stoichiometry Multiple Choice Answers:

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