

**WORKSHEET**  
**CHEMISTRY FOR ENGINEERING STUDENTS, LECTURE 1**  
 CHEMISTRY OF THE ELEMENTS OF GROUP 1: ALKALI METALS

**ALKALI METALS**

**QUESTION 1. Write the chemical equations for the reactions and products.**

1. Decomposition of sodium hydroxide used for preparing sodium carbonate which is used industrially.  $2NaOH \rightarrow Na_2CO_3 + H_2O$  (this reaction occurs at 300°C in presence of  $CaO$  catalyst)
  - (a) write the balanced chemical equation for the reaction  
 $2NaOH \rightarrow Na_2CO_3 + H_2O$
  - (b) calculate the mass of sodium carbonate and water formed from 100g of sodium hydroxide  
 $100g \text{ NaOH} \rightarrow 106g \text{ Na}_2CO_3 + 18g \text{ H}_2O$
  - (c) calculate mass of the gas evolved from 100g of sodium hydroxide  
 $18g \text{ H}_2O$
  
2. Write chemical equations involved in various reactions for sodium the alkali metal (use necessary conditions). The hydroxide reacts with acid to form salt and water and with water to form alkali and hydrogen. These three types of reactions are indicated by 1, 2 and 3 respectively.
  - (a) write the balanced chemical equation for the reaction of sodium hydroxide with hydrochloric acid  
 $NaOH + HCl \rightarrow NaCl + H_2O$
  - (b) calculate the volume of hydrogen evolved from 100g of sodium hydroxide with water  
 $100g \text{ NaOH} \rightarrow 11.2 \text{ L H}_2$
  - (c) calculate volume of carbon dioxide evolved in 100g of sodium hydroxide  
 $100g \text{ NaOH} \rightarrow 11.2 \text{ L CO}_2$
  
3. Write chemical equation for the reaction of sodium with water. Identify gas evolved in the reaction.  $2Na + 2H_2O \rightarrow 2NaOH + H_2$ 
  - (a) write the balanced chemical equation for the reaction of sodium with water  
 $2Na + 2H_2O \rightarrow 2NaOH + H_2$
  - (b) how much mass of the gas evolved by decomposition of 1 mole of sodium  
 $2g \text{ Na} \rightarrow 1g \text{ H}_2$
  - (c) how much volume of the gas evolved from 1 mole of sodium  
 $2g \text{ Na} \rightarrow 11.2 \text{ L H}_2$

**QUESTION 2. Write the chemical equations for the reactions.**

1. When metal potassium reacts with oxygen the following reaction is observed (potassium reacting with oxygen to form potassium peroxide).  $2K + O_2 \rightarrow K_2O_2$ 
  - (a) write the balanced chemical equation for the reaction  
 $2K + O_2 \rightarrow K_2O_2$
  - (b) amount of the reactants in the decomposition of  $K_2O_2$

# Chem 21 Lab Answers Byboomore

**IM Harris**



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