



Chemical Process Calculations

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Instructor:

Dr. Sushant Kumar
Department of Chemical and Biochemical Engineering
Indian Institute of Technology, Patna

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Energy Balances

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Laws of Thermochemistry:

At a given temperature and pressure, the quantity of energy required to decompose a chemical compound into its elements is precisely equal to that evolved in the formation of that compound from its elements.

Law of Constant-heat Summation:

It states that the net heat evolved or absorbed in a chemical process is the same whether the reaction takes place in one or in several steps. The total change in enthalpy of a system is dependent on the temperature, pressure, state of aggregation, and state of combination at the beginning and at the end of the reaction and is independent of the number of intermediate chemical reactions involved. This principle is known as the **law of Hess**.

Hence, using this principle, it is possible to calculate the heat of formation of a compound from a series of reactions not involving the direct formation of the compound from the elements. The majority of chemical compounds cannot be formed in the pure state directly from the

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