

Purdue University – Department of Chemical Engineering

Final Examination CHE 348
Closed book and notes

Name: _____
Friday, May 2, 2008

Solutions

1. (15 points) Decomposition of ozone catalyzed by atomic chlorine

The following two-step cycle is believed to occur for the destruction of ozone by chlorine in the stratosphere:



The overall reaction is $\text{O} + \text{O}_3 \rightarrow 2\text{O}_2$. Note that Cl and ClO are very reactive species and that Cl is regenerated at the end of the cycle (Cl is a catalyst for this reaction).

The reaction constants measured experimentally are:

$$k_1 = 5 \times 10^{-11} \exp(-140/T) \text{ cm}^3 \text{ s}^{-1} \text{ and } k_2 = 1.1 \times 10^{-10} \exp(-220/T) \text{ cm}^3 \text{ s}^{-1}$$

- Using the steady-state approximation and the fact that the number of active centers ($[\text{L}] = [\text{Cl}] + [\text{ClO}]$) is a constant show that the rate of ozone decomposition is $r_c = - (k_1 k_2 [\text{L}] [\text{O}] [\text{O}_3]) / (k_1 [\text{O}_3] + k_2 [\text{O}])$
- Since $[\text{O}] \ll [\text{O}_3]$, simplify the denominator for the expression in a). The non-catalyzed rate of ozone decomposition (r_{nc}) is elementary ($\text{O} + \text{O}_3 \rightarrow 2\text{O}_2$) with $k = 1.9 \times 10^{-11} \exp(-2300/T) \text{ cm}^3 \text{ s}^{-1}$

Calculate the ratio of the catalyzed to the non-catalyzed reaction (r_c/r_{nc}) as a function of temperature using $[\text{L}]/[\text{O}_3] = 10^{-5}$, which is the value measured in the stratosphere. Calculate the ratio r_c/r_{nc} for a temperature in the ozone layer of 200 K. Does that explain why Cl destroys the ozone layer? (remember that Cl is not consumed in the cycle, it catalyzes the reaction indefinitely)

Chemical Reaction Engineering Final Exam Solution

L Cohen



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Table of Contents Chemical Reaction Engineering Final Exam Solution

1. Understanding the eBook Chemical Reaction Engineering Final Exam Solution
 - The Rise of Digital Reading Chemical Reaction Engineering Final Exam Solution
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Reaction Engineering Final Exam Solution
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Reaction Engineering Final Exam Solution
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Reaction Engineering Final Exam Solution

- Personalized Recommendations
 - Chemical Reaction Engineering Final Exam Solution User Reviews and Ratings
 - Chemical Reaction Engineering Final Exam Solution and Bestseller Lists
5. Accessing Chemical Reaction Engineering Final Exam Solution Free and Paid eBooks
 - Chemical Reaction Engineering Final Exam Solution Public Domain eBooks
 - Chemical Reaction Engineering Final Exam Solution eBook Subscription Services
 - Chemical Reaction Engineering Final Exam Solution Budget-Friendly Options
 6. Navigating Chemical Reaction Engineering Final Exam Solution eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemical Reaction Engineering Final Exam Solution Compatibility with Devices
 - Chemical Reaction Engineering Final Exam Solution Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Reaction Engineering Final Exam Solution
 - Highlighting and Note-Taking Chemical Reaction Engineering Final Exam Solution
 - Interactive Elements Chemical Reaction Engineering Final Exam Solution
 8. Staying Engaged with Chemical Reaction Engineering Final Exam Solution
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Reaction Engineering Final Exam Solution
 9. Balancing eBooks and Physical Books Chemical Reaction Engineering Final Exam Solution
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Reaction Engineering Final Exam Solution
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemical Reaction Engineering Final Exam Solution
 - Setting Reading Goals Chemical Reaction Engineering Final Exam Solution
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemical Reaction Engineering Final Exam Solution

- Fact-Checking eBook Content of Chemical Reaction Engineering Final Exam Solution
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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