

MOST EXPECTED CHEMISTRY QUESTIONS

UNIT : 1 : SOLUTIONS (7 MARKS)

1. Define: Molality, Molarity, Mass percentage, Volume percentage, Parts per million (ppm), Mole fraction. Write their formulas also.
2. How does a change in temperature influence values of molarity and molality.
3. State Henry law with its mathematical expressions. What is the significance of Henry's Law constant K_H ? Mention some of important applications Henry law.
4. Why do aquatic species remain more comfortable in lakes in winters than in summers?
5. State Raoult's law for a solution of volatile liquids. Give its mathematical relationship.
6. What is an ideal solution? What type of solutions are likely to behave as ideal solutions? Draw the plot of vapour pressure and mole fraction of an ideal solution at constant temperature.
7. Differentiate between Non ideal solutions exhibiting Positive deviations & Negative deviations
8. What are Azeotropes? Give one example each of minimum boiling and maximum boiling azeotropes.
9. Define. (a) osmosis (b) osmotic pressure (c) reverse osmosis & application (d) colligative properties
10. Measurement of osmotic pressure method is preferred for the determination of molecular masses of macromolecules such as proteins and polymers. Give two reasons.
11. What will happen if RBC are placed in (i) 0.5% NaCl Solution (ii) 1% NaCl Solution?
12. Define *abnormal molecular mass*. What is Van't Hoff's factor? What is value of Van't Hoff's factor when the solute undergoes (a) association (b) dissociation?
13. Calculate the molarity of a solution containing 5 g of NaOH in 450 ml solution.
14. Concentrated nitric acid used in the laboratory work is 68% nitric acid by mass in aqueous solution. What should be the molarity of such a sample of the acid if the density of the solution is 1.504 g/ml?
15. Calculate molality of 2.5 g of ethanoic acid (CH_3COOH) in 75 g of benzene.
16. An antifreeze solution is prepared from 222.6 g of ethylene glycol, $\text{C}_2\text{H}_4(\text{OH})_2$ and 200 g of water. Calculate molality of solution. If the density of the solution is 1.072 g/ml then what shall be the molarity of the solution?
17. A solution of glucose in water is labelled as 10% w/w. What should be the molality and mole fraction of each component in the solution? If the density of solution is 1.2 g/ml, then what shall be the molarity of the solution?
18. If N_2 gas is bubbled through water at 298 K, how many millimoles of N_2 gas would dissolve in 1 litre of water. Assume that N_2 exerts a partial pressure of 0.987 bar. Henry's law constant for N_2 at 293 K is 76.48 bar.
19. Vapour pressure of chloroform (CHCl_3) and dichloroform (CH_2Cl_2) at 298 K are 200 mm Hg and 415 mm Hg respectively. (i) Calculate the vapour pressure of the solution prepared by mixing 25.5 g of CHCl_3 and 40.0 g of CH_2Cl_2 at 298 K and (ii) mole fractions of each component in vapour phase.
20. Vapour pressure of water at 293 K is 17.535 mm Hg. Calculate vapour pressure of water at 293 K when 25 g of glucose is dissolved in 450 g of water.
21. A solution is prepared by dissolving 10 g of non volatile solute in 200 g of water. It has a vapour pressure of 31.84 mm of Hg at 308 K. Calculate the molar mass of the solute. (Vapour pressure of pure water at 308 K is 32 mm of Hg.)
22. 18 g glucose $\text{C}_6\text{H}_{12}\text{O}_6$ is dissolved in 1 kg of water in a saucepan. At what temperature will solution boil? K_b for water is $0.512 \text{ K kg mol}^{-1}$.
23. A solution of glycerol ($\text{C}_3\text{H}_8\text{O}_3$) in water was prepared by dissolving some glycerol in 500 g of water. This solution has a boiling point of 100.42°C . What mass of glycerol was dissolved to make this solution? (K_b for water = $0.512 \text{ K kg mol}^{-1}$)

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