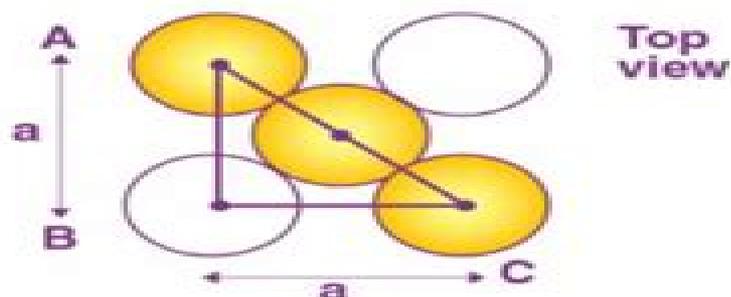


Let the edge length of the unit cell = a
 let the radius of each sphere = r
 Thus, $AC = 4r$
 From the right angled triangle ABC , we have:
 $AC = (a^2 + a^2)^{1/2} = a(2)^{1/2}$



Therefore, $4r = a(2)^{1/2}$
 $\Rightarrow a = 4r / (2)^{1/2}$
 Thus, Volume of unit cell $= a^3 = \{ 4r / (2)^{1/2} \}^3$
 $a^3 = 64r^3 / 2(2)^{1/2} = 32r^3 / (2)^{1/2}$
 No. of unit cell in FCC = 4
 Volume of four spheres = $4 \times \pi r^3 (4/3)$
 Thus, packing efficiency = $[\pi r^3 (16/3)] / [32r^3 / (2)^{1/2}] = 0.74$ or 74 %

Q9. Silver crystallises in fcc lattice. If edge length of the cell is 4.07×10^{-8} cm and density is 10.5 g cm^{-3} , calculate the atomic mass of silver.

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Table of Contents Chemistry 12 Solid State Ncert Solutions

1. Understanding the eBook Chemistry 12 Solid State Ncert Solutions
 - The Rise of Digital Reading Chemistry 12 Solid State Ncert Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 12 Solid State Ncert Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 12 Solid State Ncert Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 12 Solid State Ncert Solutions
 - Personalized Recommendations

-
- Chemistry 12 Solid State Ncert Solutions User Reviews and Ratings
 - Chemistry 12 Solid State Ncert Solutions and Bestseller Lists
5. Accessing Chemistry 12 Solid State Ncert Solutions Free and Paid eBooks
 - Chemistry 12 Solid State Ncert Solutions Public Domain eBooks
 - Chemistry 12 Solid State Ncert Solutions eBook Subscription Services
 - Chemistry 12 Solid State Ncert Solutions Budget-Friendly Options
 6. Navigating Chemistry 12 Solid State Ncert Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry 12 Solid State Ncert Solutions Compatibility with Devices
 - Chemistry 12 Solid State Ncert Solutions Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry 12 Solid State Ncert Solutions
 - Highlighting and Note-Taking Chemistry 12 Solid State Ncert Solutions
 - Interactive Elements Chemistry 12 Solid State Ncert Solutions
 8. Staying Engaged with Chemistry 12 Solid State Ncert Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 12 Solid State Ncert Solutions
 9. Balancing eBooks and Physical Books Chemistry 12 Solid State Ncert Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 12 Solid State Ncert Solutions
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry 12 Solid State Ncert Solutions
 - Setting Reading Goals Chemistry 12 Solid State Ncert Solutions
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry 12 Solid State Ncert Solutions
 - Fact-Checking eBook Content of Chemistry 12 Solid State Ncert Solutions

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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