

## 10

## CHEMICAL QUANTITIES

## Practice Problems

On a piece of paper, solve the following problems.

**YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!**

## SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER

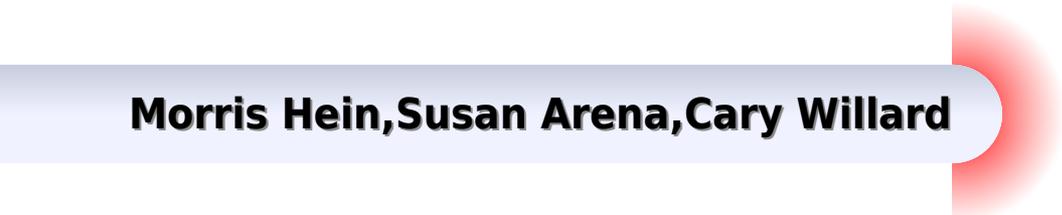
1. What is the molar mass of sucrose ( $C_{12}H_{22}O_{11}$ )?
2. What is the molar mass of each of the following compounds?
  - a. phosphorus pentachloride ( $PCl_5$ )
  - b. uranium hexafluoride ( $UF_6$ )
3. Calculate the molar mass of each of the following ionic compounds:
  - a.  $KMnO_4$
  - b.  $Ca_3(PO_4)_2$
4. How many moles is  $3.52 \times 10^{24}$  molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen ( $O_2$ )?

## SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS

1. What is the molar mass of each of the following compounds?
  - a.  $C_6H_{12}O_6$
  - b.  $NaHCO_3$
  - c.  $C_2H_{12}$
  - d.  $KNH_4SO_4$
2. Calculate the mass in grams of each of the following:
  - a. 8.0 mol lead oxide ( $PbO$ )
  - b. 0.75 mol hydrogen sulfide ( $H_2S$ )
  - c. 0.00100 mol silicon tetrahydride ( $SiH_4$ )
  - d.  $1.50 \times 10^{-2}$  mol molecular oxygen ( $O_2$ )
  - e. 2.30 mol ethylene glycol ( $C_2H_6O_2$ )
3. How many grams are in 1.73 mol of dinitrogen pentoxide ( $N_2O_5$ )?
4. How many grams are in 0.658 mol of calcium phosphate [ $Ca_3(PO_4)_2$ ]?
5. Calculate the number of moles in each of the following:
  - a. 0.50 g sodium bromide ( $NaBr$ )
  - b. 13.5 g magnesium nitrate [ $Mg(NO_3)_2$ ]
  - c. 1.02 g magnesium chloride ( $MgCl_2$ )
  - d. 0.00100 g monochloromethane ( $CH_3Cl$ )
  - e.  $1.50 \times 10^{-3}$  g propylene glycol [ $C_3H_8(OH)_2$ ]
6. A chemist plans to use 435.0 grams of ammonium nitrate ( $NH_4NO_3$ ) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ( $C_{20}H_{24}N_2O_2$ ). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane ( $CH_4$ ) gas?
9. How many moles is 135 L of ammonia ( $NH_3$ ) gas at STP?

# Chemistry Chapter 10 The Mole Study Answers

**Morris Hein, Susan Arena, Cary Willard**



## **Chemistry Chapter 10 The Mole Study Answers:**

## Reviewing **Chemistry Chapter 10 The Mole Study Answers**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Chemistry Chapter 10 The Mole Study Answers**," an enthralling opus penned by a highly acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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