

## Chemistry - Chapter 7 Study Guide

### Section 7.1:

- *Valence electrons* – electrons in the outer energy level
  - responsible for bonding with other atoms
- How do you know how many valence electrons an element has?
  - Look at the group number
    - Group 1A = 1 valence electron
    - Group 7A = 7 valence electrons
    - Etc...
- *Electron Dot Diagrams* – a way of showing and keeping track of valence electrons
  - Written using the following steps:
    1. Write the element symbol – representing the nucleus and inner electrons
    2. Put one dot for each valence electrons,
    3. Don't forget to add them using Hund's rule – they don't pair up unless they have to.
- *The Octet Rule* – atoms tend to want to achieve a noble gas configuration because it is the most stable. Octet = eight electrons
- *Cations* – a positive ion (usually metals)
  - Lose electrons to attain a noble gas configuration
  - Written like  $\text{Na}^{1+}$
  - No dots are shown (on an electron dot diagram) because they lose the valence electrons they had
- *Anions* – a negative ion (usually non-metals)
  - Gain electrons to attain a noble gas configuration
  - Written like  $\text{P}^{3-}$
  - Dots are shown because valence electrons are gained
- 8 valence electrons = stable atom (noble gases.)

### Section 7.2:

- Anions (-) and cations (+) are held together by opposite charges; a bond is formed through the transfer of electrons to achieve a noble gas configuration (octet.)
- *Ionic Compounds:*
  - Also called salts
  - Made from a cation and an anion
    - Which means they're made of a metal and a non-metal
- **\*if you don't understand how to do ionic bonding, check the book or ask somebody to help... it's hard to explain\***
- The result of ionic bonding is an ionic compound, shown in a *formula unit*
  - Written like  $\text{Ca}_3\text{P}_2$

# Chemistry Matter And Change Chapter 7 Study Guide Answer Key

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## **Table of Contents Chemistry Matter And Change Chapter 7 Study Guide Answer Key**

1. Understanding the eBook Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - The Rise of Digital Reading Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - Personalized Recommendations
  - Chemistry Matter And Change Chapter 7 Study Guide Answer Key User Reviews and Ratings
  - Chemistry Matter And Change Chapter 7 Study Guide Answer Key and Bestseller Lists

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  - Chemistry Matter And Change Chapter 7 Study Guide Answer Key Budget-Friendly Options
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  - Chemistry Matter And Change Chapter 7 Study Guide Answer Key Enhanced eBook Features
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  - Highlighting and Note-Taking Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - Interactive Elements Chemistry Matter And Change Chapter 7 Study Guide Answer Key
8. Staying Engaged with Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Matter And Change Chapter 7 Study Guide Answer Key
9. Balancing eBooks and Physical Books Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Matter And Change Chapter 7 Study Guide Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - Setting Reading Goals Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - Fact-Checking eBook Content of Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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