

# Chapter 9 Notes: Stoichiometry & the Mole

Chemistry Honors/C101/C121

## Mole Conversions with Moleville Map

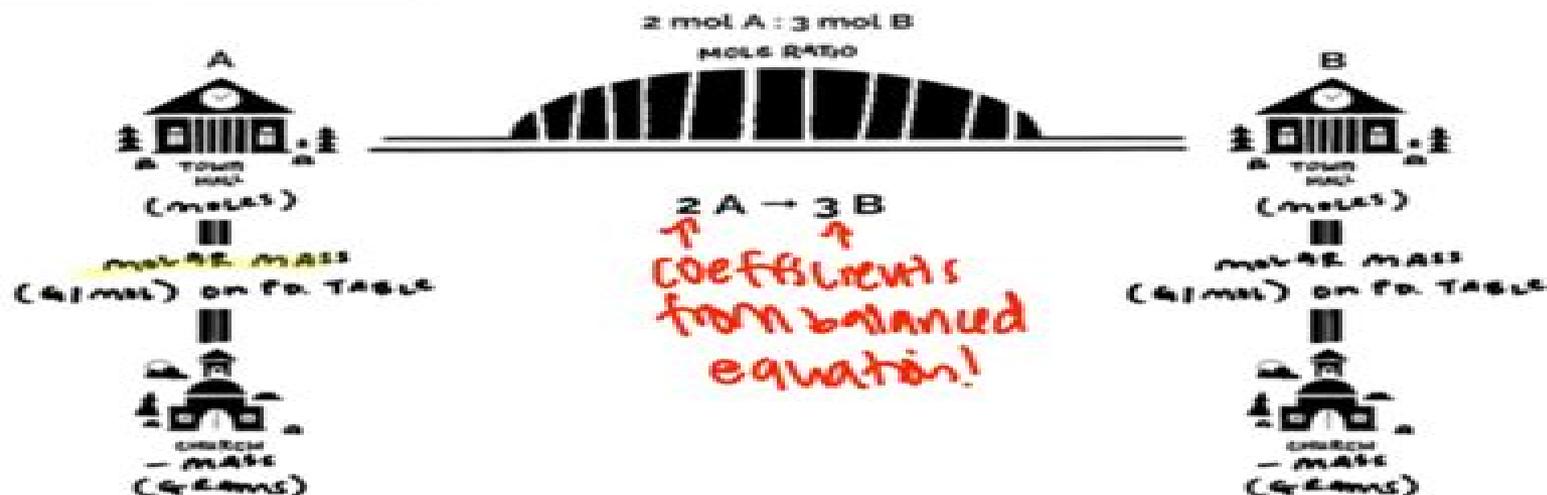
**Given:**  
Avogadro's Number  
• 1 mole =  $6.022 \times 10^{23}$  particles (atoms/molecules/compounds)

**Know:**  
Molar Mass  
• 1 mole = [mass on PT] grams

## Stoichiometry

- The study of the relationships between the amount of reactants used and amount of products formed by a chemical reaction.
- Etymology from Greek:
  - stoikheion = "element"
  - metron = "measure"
- Based on the law of conservation of mass, which states that matter can neither be created nor destroyed in a chemical reaction.

## Moleville A to Moleville B



# Chemistry Notes Chapter 9 Stoichiometry

**Sebastian Brünink**



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