

Chapter 6 Thermochemistry

Homework Problems

- The enthalpy change for the fermentation (without oxygen, O_2 gas) reaction of glucose to alcohol and carbon dioxide is -67 kJ.
 - Write a chemical equation for the fermentation reaction and show how energy is involved in the reaction.
 - Is the reaction exothermic or endothermic?
 - Is energy, in the form of heat, absorbed or evolved as the reaction occurs?
- The specific heat capacity of graphite is 0.71 J/°C·g. Calculate the energy required to:
 - Raise the temperature of 1.0 mol graphite by 1.0 °C.
 - Raise the temperature of 850 g graphite by 150 °C.
 - Raise the temperature of 75 kg graphite from 294 K to 348 K.
- In a coffee-cup calorimeter, 100.0 mL of 1.0 M NaOH and 100.0 mL of 1.0 M HCl are mixed. Both solutions were originally at 24.6 °C. After the reaction, the final temperature is 31.3 °C. Assume that all the solutions have a density of 1.0 g/mL and a specific heat capacity of 4.18 J/g°C. Calculate the enthalpy change for the neutralization of HCl by NaOH. Assume that no heat is lost to the surroundings or to the calorimeter.
- Camphor ($C_{15}H_{26}O$) has an energy of combustion of -5903.6 kJ/mol. A sample of camphor with a mass of 0.1204 g is burned in a bomb calorimeter, the temperature increases by 2.28 °C. Calculate the heat capacity of the calorimeter.
A 0.1964 g sample of quinone ($C_9H_6O_2$) is burned in a bomb calorimeter that has a heat capacity of 1.56 kJ/°C. The temperature of the calorimeter increases by 3.2 °C. Calculate the energy of combustion of quinone per gram and per mole.

Chapter 6 Thermochemistry Homework Problems

Jiyuan Zhang



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and equipment Practice drills at the end of each content review chapter plus step by step walk throughs of sample exam questions

Solutions Manual to Accompany General Chemistry with Qualitative Analysis, Second Edition Ralph H. Petrucci, Robert K. Wismer, 1987

Chemistry Steven S. Zumdahl, Kenneth C. Brooks, 1986

Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg, 2015-01-16

Ebook Chemistry The Molecular Nature of Matter and Change **CHEMISTRY SILBERBERG**, 2003

Rocket Propulsion Stephen D. Heister, William E. Anderson, Timothée L. Pourpoint, R. Joseph Cassady, 2019-02-07

A modern pedagogical treatment of the latest industry trends in rocket propulsion developed from the authors extensive experience in both industry and academia Students are guided along a step by step journey through modern rocket propulsion beginning with the historical context and an introduction to top level performance measures and progressing on to in depth discussions of the chemical aspects of fluid flow combustion thermochemistry and chemical equilibrium solid liquid and hybrid rocket propellants mission requirements and an overview of electric propulsion

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Schaum's Outline of Theory and Problems of College Chemistry Jerome Laib Rosenberg, Lawrence M. Epstein, 1990

Introduction to Supercritical Fluids Richard Smith, Hiroshi Inomata, Cor Peters, 2013-12-08

This text provides an introduction to supercritical fluids with easy to use Excel spreadsheets suitable for both specialized discipline chemistry or chemical engineering student and mixed discipline engineering economic student classes Each chapter contains worked examples tip boxes and end of the chapter problems and projects Part I covers web based chemical information resources applications and simplified theory presented in a way that allows students of all disciplines to delve into the properties of supercritical fluids and to design energy extraction and materials formation systems for real world processes that use supercritical water or supercritical carbon dioxide Part II takes a practical approach and addresses the thermodynamic framework equations of state fluid phase equilibria heat and mass transfer chemical equilibria and reaction kinetics of supercritical fluids

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Study Guide [to Accompany] General Chemistry Dixie Goss, 2002

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Physical Chemistry, Series One: Thermochemistry and thermodynamics, edited by H. A. Skinner Amyand David Buckingham,1972 Chemical Principles William L. Masterton,Emil J. Slowinski,Conrad L. Stanitski,1981 *Engineering News* ,1906 General Chemistry Key Concise Review Guide Sixth Edition Darrell D. Ebbing,1999

Complete Preparation for the MCAT Williams & Wilkins Review,1998-04 Here is the most respected test prep book for the Medical College Admission Test you can buy featuring an active learning approach for a better understanding of the exam s content and a better chance for success Unique to this guide are coverage of all recent changes in the MCAT plus a step by step plan for sharpening cognitive skills developing problem solving skills and critical thinking This thorough guide replaces expensive test preparation courses while giving students exactly what they need to get ready for the MCAT

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