

CHAPTER 4 The Structure of Matter

SECTION

2

Ionic and Covalent Bonding

KEY IDEAS

As you read this section, keep these questions in mind:

- Why do atoms form bonds?
- How do ionic bonds and covalent bonds differ?
- What gives metals their distinctive properties?
- What makes a polyatomic ion different from other ions?

Why Do Chemical Bonds Form?

Atoms form bonds when their valence electrons interact. Recall that atoms with filled outermost energy levels are more stable than atoms with partially filled energy levels. In general, atoms join to form bonds so that each atom has a stable electron configuration. In other words, each atom binds in order to fill its outermost energy levels. ☐

There are two basic kinds of chemical bonding: ionic bonding and covalent bonding. The type of bonding in a compound determines many of the properties of the compound.

A Comparison of Ionic and Covalent Compounds		
	ionic compounds	covalent compounds
Structure	network of bonded ions	molecules
Valence electrons	transferred	shared
Electrical conductivity	good (when melted or dissolved)	poor
State at room temperature	solid	solid, liquid, or gas
Melting and boiling points	generally high	generally low

What Are the Properties of Ionic Bonds?

Ionic bonds form between oppositely charged ions. In general, atoms of metals, such as sodium and calcium, form positively charged ions. Atoms of nonmetals, such as chlorine and oxygen, form negatively charged ions. The attraction between ions with opposite charges holds ionic compounds together. ☐

READING TOOLBOX

Compare As you read, make a chart that describes and compares the different types of bonds.

READING CHECK

1. **Explain** Why do atoms join to form bonds?

LOOKING CLOSER

2. **Compare** How do the structures of ionic and covalent compounds differ?

READING CHECK

3. **Identify** What holds ionic compounds together?

Chapter Section 2 Ionic And Covalent Bonding

Rosina Ehmman



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