

## Chapter-1: Safe Use of Laboratory

Ques. ► 1

Na

Chemical-A

H<sub>2</sub>SO<sub>4</sub>

Chemical-B

(R.B., C.B., C.C.B., B.B.-15)

- What is suspension?
- Why the 19<sup>th</sup> electron of potassium goes to 4s orbital instead of 3d?
- Explain the safe preservation of chemical-A in the laboratory.
- "Optimum use of chemical-B in the lab is important for conserving environment"- Analyze it/Justify.

**Answer to the question no. 1**

**a.** When a substance is dispersed in another substance as particle having radius greater than 10.5  $\mu\text{m}$  forming a heterogeneous and temporary mixture is known as suspension.

**b.** Electronic configuration of potassium is :



According to Aufbau principle, the orbital having the least (n + l) value is the orbital with least energy and electron will first enter this orbital.

For 3d orbital, n = 3, l = 2 ; n + l = 5

For 4s orbital, n = 4, l = 0 ; n + l = 4

So for K, 4s orbital has lower energy than 3d orbital. For this reason, the 19<sup>th</sup> electron of K enters 4s orbital instead of 3d orbital following Aufbau principle.

**c.** The chemical substance A is metallic sodium (Na).

Na is a reactive metal with instantaneous interaction with water. These type of substance react with water and form combustible hydrogen gas. So they are preserved in a cool, dry place so that they don't come in contact with water and air. For this reason, metallic Na is kept in kerosene or mineral oil. If Na is kept free, it will come in contact with moist air and will form sodium oxide which will then react with water vapour to form sodium hydroxide. Na is a reactive metal. Na reacts with O<sub>2</sub> of air. Na has no reaction with hydrocarbon. It is heavier than most hydrocarbons. Kerosene is a non toxic liquid and it doesn't vaporise easily. Na also reacts with Alcohol.

It will again react with the CO<sub>2</sub> of air and will produce sodium carbonate. As a result Na metal loses it's elementary characteristics.



**d.** The chemical-B of the stem is H<sub>2</sub>SO<sub>4</sub>. The significance of optimum use of chemical substances is enormous.

Based on it's nature, almost every chemical substance has more or less health and environmental impact. In this case, the more

is the contact the more will be the impact. So it is important to reduce the hazardous impact of chemical substances on health and environment and to ensure the economic use of it. It is required to have a plan before each that to reduce the use of chemical at a minimum.

In higher secondary level, the tests which will give accurate result if done in groups, will save huge amount of chemicals and will reduce the hazardous impact on the students and environment. The importance of optimum use of chemical substances in the lab are given below:

- In smoke test, using excess sample will produce excess smoke and thus increases the chances of harming the students and environment.
- Using excess reactant will increase the need for more reagent thus increasing the cost of experiment. Furthermore, there will be excess chemical waste thus creating complexity in waste management and disposal.
- Optimum use of chemical will reduce the risk on health and environment.
- If we use less amount of harmful and poisonous chemicals, we will require less protective measures. As a result our experiments will be easier.
- Use of excess combustible substances increase the chances of fire.
- Finishing the experiment with less reactant will require less heat application. It will reduce fuel cost and the environmental pollution will be reduced.

**Ques. ► 2** The substances which are used carefully in the laboratory are —

- Concentrated Sulfuric acid
- Liquid Ammonia
- Sodium metal
- Potassium Ferrocyanide

(R.B., C.B., C.C.B., B.B.-15)

- What is chromatography?
- Which one between Na<sup>+</sup> and F<sup>-</sup> is big in size?
- Why is not water directly added to the chemicals (i) and (iii) of the stem? Explain.
- Analyze the impact on environment for the excess uses of the chemicals of the stem.

**Answer to the question no. 2**

**a.** The process of separating a mixture into it's component by panning it through a static phase using a mobile phase of either liquid or gas. Where separation occurs based on different rate of adsorption or different solubility and distribution coefficient is known as chromatography.

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