

# STOICHIOMETRY PROBLEMS

1. The combustion of a sample of butane,  $C_4H_{10}$  (lighter fluid), produced 2.46 grams of water.



- (a) How many moles of water formed?
- (b) How many moles of butane burned?
- (c) How many grams of butane burned?
- (d) How much oxygen was used up in moles?
- (e) How much oxygen was used up in grams?

2. Sodium metal burns in air according to the balanced reaction below.



- (a) How many moles of oxygen are needed to react with 9.5 g of sodium completely?
- (b) How many grams of sodium are needed to produce 12.5 g of sodium oxide?

3. In the combustion of 54.5 g of butane ( $C_4H_{10}$ ), how many grams of  $CO_2$  are produced? Write and balance the equation before solving.

# Guided Practice Problems 11 Stoichiometry

**Brent L. Iverson**



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