

## Chemistry Periodic Table Test Chapter 6 Chem IH

**Matching** Match each item with the correct statement below.

- |                      |                     |
|----------------------|---------------------|
| a. electronegativity | e. transition metal |
| b. ionization energy | f. periodic law     |
| c. electrons         | g. group            |
| d. metal             | h. period           |

- \_\_\_\_\_ 1. vertical column in the periodic table  
\_\_\_\_\_ 2. type of element characterized by the presence of electrons in the *d* orbital  
\_\_\_\_\_ 3. type of element characterized by losing electrons easily  
\_\_\_\_\_ 4. ability of an atom to attract electrons when the atom is in a compound  
\_\_\_\_\_ 5. energy required to remove an electron from an atom

### Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 6. What is another name for the main group elements?  
a. s block elements  
b. p block elements elements  
c. s & p block elements  
d. d block elements
- \_\_\_\_\_ 7. What is another name for the transition metals?  
a. noble gases  
b. s block elements elements  
c. p block elements elements  
d. d block elements
- \_\_\_\_\_ 8. Who arranged the elements according to atomic mass and used the arrangement to predict the properties of missing elements?  
a. Henry Moseley  
b. Antoine Lavoisier  
c. John Dalton  
d. Dmitri Mendeleev
- \_\_\_\_\_ 9. Which of the following categories includes the majority of the elements?  
a. metalloids  
b. liquids  
c. metals  
d. nonmetals
- \_\_\_\_\_ 10. In which of the following sets is the symbol of the element, the number of protons, and the number of electrons given correctly?  
a. In, 49 protons, 49 electrons  
b. Zn, 30 protons, 60 electrons  
c. Cs, 55 protons, 132.9 electrons  
d. F, 19 protons, 19 electrons
- \_\_\_\_\_ 11. The atomic number of an element is the total number of which particles in the nucleus?  
a. neutrons  
b. protons  
c. electrons  
d. protons and electrons
- \_\_\_\_\_ 12. What element has the electron configuration  $1s^2 2s^2 2p^6 3s^2 3p^2$ ?  
a. nitrogen  
b. selenium  
c. silicon  
d. silver
- \_\_\_\_\_ 13. Which subatomic particle plays the greatest part in determining the properties of an element?  
a. proton  
b. electron  
c. neutron  
d. none of the above
- \_\_\_\_\_ 14. Which of the following elements is a transition metal?  
a. cesium  
b. copper  
c. tellurium  
d. tin

# Periodic Table Chapter Test A Answer Key

**M Woodhall**



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