

Question 1

1(a) Balance (top loading), spatula, stirring rod, dropper, pipette (10 cm³), pipette (25 cm³), glass (filter) funnel, burette.

1(b)

Burette readings (cm ³)	Titration 1	Titration 2	Titration 3
Final volume	2.00	1.85	1.50
Initial volume	21.95	21.85	19.50
Volume of solution used	19.95*	20.00*	18.00

1(c)(i) See table above.

1(c)(ii) Average volume = (volume used in titration 1 + volume used in titration 2)/2.

Average volume = $\{(19.95 + 20.00)/2\}$ cm³. i.e. 19.98 cm³.

Actual accurate value to be used in subsequent calculations is 19.975 cm³.

1(d) 1000 cm³ K₂Cr₂O_{7(aq)} contains 0.020 mol. K₂Cr₂O₇. 19.98 cm³ K₂Cr₂O_{7(aq)} contains $\{19.98 \times (0.020/1000)\}$ mol. K₂Cr₂O₇, i.e. 4.00×10^{-4} mol K₂Cr₂O₇.

Actual accurate value to be used in subsequent calculations is 3.995×10^{-4} mol K₂Cr₂O₇.

1(e) From the balanced equation, the Fe²⁺:Cr₂O₇²⁻ stoichiometric ratio is 6:1. So 6 moles of iron ions (Fe²⁺) react with one mole of dichromate ions (Cr₂O₇²⁻).

1(f) Moles of iron in sample = (6 × moles of Cr₂O₇²⁻ used in titration) i.e. $(6 \times 4.00 \times 10^{-4})$ mol. = 2.4×10^{-3} mol. iron.

Actual accurate value to be used in subsequent calculations is 2.397×10^{-3} mol.

1(g) Mass of iron in salt sample = (moles of iron in sample × molar mass of iron).

Mass of iron in salt sample = $(2.4 \times 10^{-3} \times 55.8)$ g iron i.e. 0.134 g iron.

Actual accurate value to be used in subsequent calculations is 0.1337526 g.

1(h) Percentage of iron in iron salt sample = (mass of iron in sample/mass of sample) × 100%.

Percentage of iron in iron salt sample = $\{0.134/0.500\} \times 100\%$ i.e. 26.80% iron.

If accurate values were used in all calculation steps leading to this one, the final answer would be 26.75%.

1(i) Reduction of a species is the gain of an electron, or electrons, by that species.

1(j) A reducing agent is a species which causes the reduction of another species, and is itself oxidized when it reacts; it loses an electron, or electrons.

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1. Understanding the eBook Cxc Chemistry Past Paper Questions
 - The Rise of Digital Reading Cxc Chemistry Past Paper Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Cxc Chemistry Past Paper Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Cxc Chemistry Past Paper Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Cxc Chemistry Past Paper Questions
 - Personalized Recommendations
 - Cxc Chemistry Past Paper Questions User Reviews and Ratings
 - Cxc Chemistry Past Paper Questions and Bestseller Lists
5. Accessing Cxc Chemistry Past Paper Questions Free and Paid eBooks
 - Cxc Chemistry Past Paper Questions Public Domain eBooks
 - Cxc Chemistry Past Paper Questions eBook Subscription Services
 - Cxc Chemistry Past Paper Questions Budget-Friendly Options
6. Navigating Cxc Chemistry Past Paper Questions eBook Formats

- ePub, PDF, MOBI, and More
 - Cxc Chemistry Past Paper Questions Compatibility with Devices
 - Cxc Chemistry Past Paper Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Cxc Chemistry Past Paper Questions
 - Highlighting and Note-Taking Cxc Chemistry Past Paper Questions
 - Interactive Elements Cxc Chemistry Past Paper Questions
 8. Staying Engaged with Cxc Chemistry Past Paper Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Cxc Chemistry Past Paper Questions
 9. Balancing eBooks and Physical Books Cxc Chemistry Past Paper Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Cxc Chemistry Past Paper Questions
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Cxc Chemistry Past Paper Questions
 - Setting Reading Goals Cxc Chemistry Past Paper Questions
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Cxc Chemistry Past Paper Questions
 - Fact-Checking eBook Content of Cxc Chemistry Past Paper Questions
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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